



$$\text{Molar ratio} = \frac{\frac{1.00}{1.00}}{\frac{3.96}{3.00}} = 0.7576$$

$$\text{wt}\% = \frac{17.1 \text{ mg} \times 175.19 \frac{\text{g}}{\text{mol}} \times 0.7576 \times 1.00}{13.5 \text{ mg} \times 168.19 \frac{\text{g}}{\text{mol}}} \times 100\% = 100\%$$

